

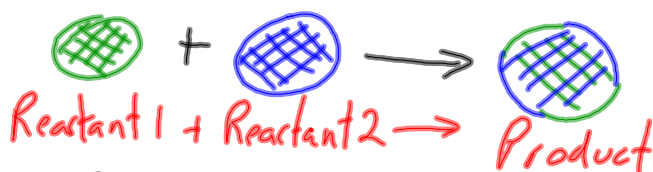
Types of Chemical Reactions

1. Synthesis-
2. Decomposition
3. Single Displacement
4. Double Displacement

Reactants \longrightarrow Products

Mar 11-12:53 PM

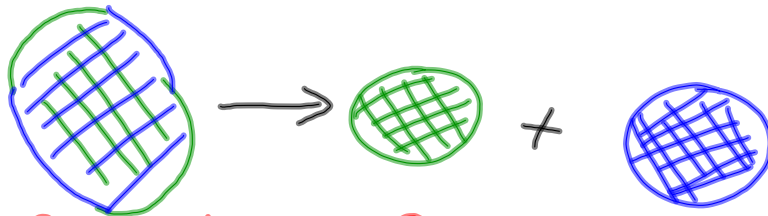
Synthesis - combination of smaller atoms and/or molecules into larger molecules.



When four hydrogen atoms combine with oxygen they produce two molecules of water

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Decomposition
splitting of large molecules
 into smaller atoms and/or
molecules



Reactant \rightarrow Product 1 + Product 2

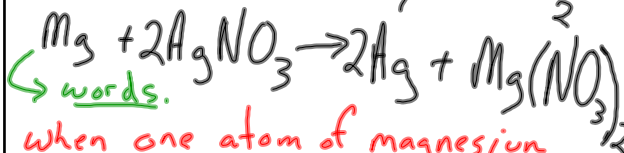


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Single Displacement
only one element replaces
another element from a
compound.



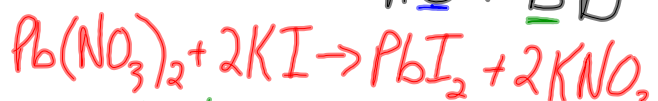
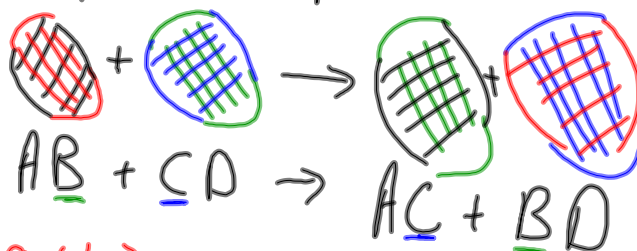
Reactant₁ + Reactant₂ \rightarrow Product₁ + Product₂



when one atom of magnesium combines with two molecules of the compound silver nitrate they together create two atoms of the element silver and one molecule of the compound magnesium nitrate.

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Double displacement
when elements in different
compounds replace each other



Lead nitrate with two molecules of
potassium iodide create lead iodide
and two molecules of potassium
nitrate.

Mar 11-1:12 PM